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problems. First, we will review and practice how to write formulas for compounds when given the compound's name. Second, we

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the formulas for the

compounds formed

from these pairs of

ions. c.  $\text{Ca}^{2+}$   $\text{K}^+$  e.  $\text{O}^{2-}$

$\text{Cu}^{+}$   $\text{Cl}^{-}$  h.  $\text{Sn}^{4+}$   $\text{Cl}^{-}$  3. Name

the following binary



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ionic compounds. a.  
MnO<sub>2</sub> b. c. CaCl<sub>2</sub> d.  
SrBr<sub>2</sub> e. f. Chapter 9  
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The formulas and  
common names for  
several substances are  
given below. 5 FePO<sub>4</sub>  
iron(III) phosphate.

Naming chemical

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The name of a base is  
the name of the cation  
followed by the name  
of the anion  
(hydroxide). The  
formula of a base is  
written by showing the

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number of hydroxide ions needed to balance the positive charge on the cation. After reading Lesson 9.4, answer the following questions. Names and Formulas of Acids hydrogen1.

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